


EELE408 Photovoltaics

Lecture 05: Semiconductor Basics

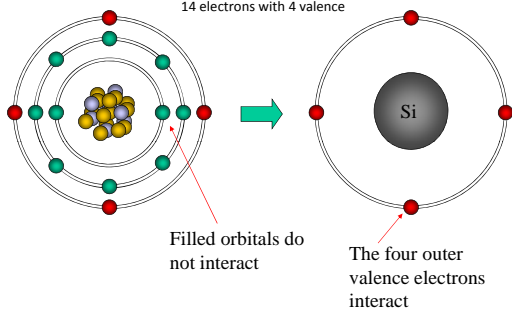
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Bond Model

Silicon Atom



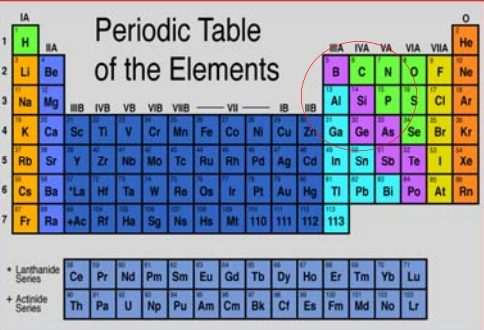
14 electrons with 4 valence

Filled orbitals do not interact

The four outer valence electrons interact

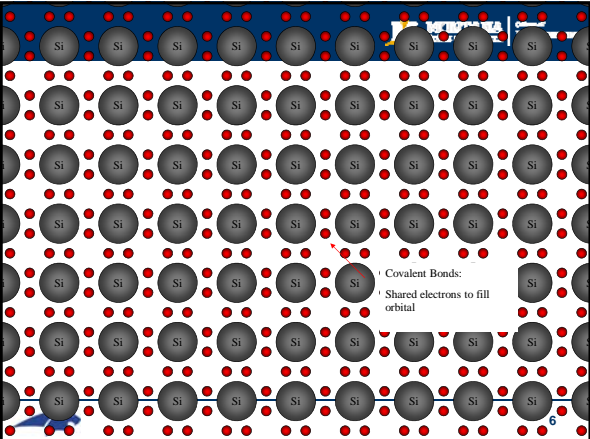
Semiconductors

Periodic Table of the Elements



Elements around Silicon

III			IV			V		
5 B Boron 10.81	6 C Carbon 12.01	7 N Nitrogen 14.01	13 Al Aluminum 26.98	14 Si Silicon 28.09	15 P Phosphorus 30.97	31 Ga Gallium 69.72	32 Ge Germanium 72.64	33 As Arsenic 74.92



Covalent Bonds:
 Shared electrons to fill orbital

Intrinsic Silicon

Poor conductor: No free electrons to carry current
Need to engineer electrical properties (conduction)

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Valence V: n-type doping

III	IV	V
5 B Boron 2.34	6 C Carbon 2.62	7 N Nitrogen 2.21
13 Al Aluminum 2.29	14 Si Silicon 2.33	15 P Phosphorus 2.42
31 Ga Gallium 2.41	32 Ge Germanium 2.32	33 As Arsenic 2.72

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Extra Negative Electrons

Each N-type dopant brings an extra electron to the lattice

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Valence III: p-type doping

III	IV	V
5 B Boron 2.34	6 C Carbon 2.62	7 N Nitrogen 2.21
13 Al Aluminum 2.29	14 Si Silicon 2.33	15 P Phosphorus 2.42
31 Ga Gallium 2.41	32 Ge Germanium 2.32	33 As Arsenic 2.72

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Extra Positive Holes

Each P-type dopant is short an electron, creating a hole in the lattice

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Bond Model

- Electrons from broken bonds are free to move
- Electrons from neighboring bonds can also move into the 'hole' created by a broken bond, allowing the broken bond or 'hole' to propagate as if had a positive charge
 - Bubble in a liquid

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Band Model

Orbital Shells

- The positions of the electrons around the nucleus are quantized to specific energy levels or **Shells**
- The orbital determines the energy of the electron

$$V = \frac{-q^2}{4\pi\epsilon_0 r}$$

n = 1
n = 2
n = 3

Energy of Electrons

Lower Electron Energy: More Tightly Bound to Nucleus

Higher Electron Energy: Less Bound to Nucleus

Orbitals fill from the bottom up.

n=1 n=2 n=3 n=4 n=5 n=6 n

Silicon Electron Configuration

[Ne]3s²3p²

Splitting of Energy Levels in a Crystalline Lattice

Continuum of Energy States

Discrete Energy States

Forbidden

Allowed

Forbidden

Band gaps

Allowed

Forbidden

Allowed and forbidden energy levels at atomic spacing

Interatomic Distance d

Band Diagram

Energy

Conduction Band

Energy Gap

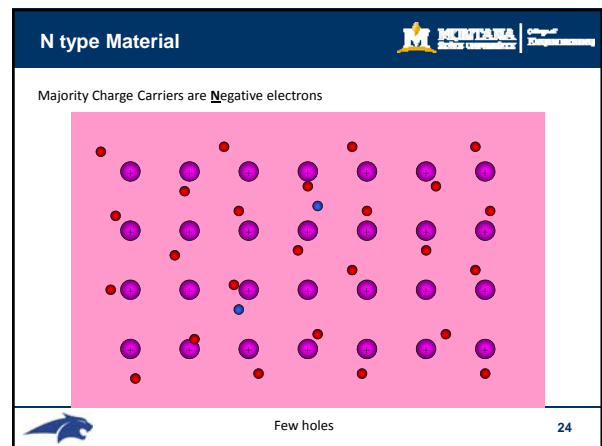
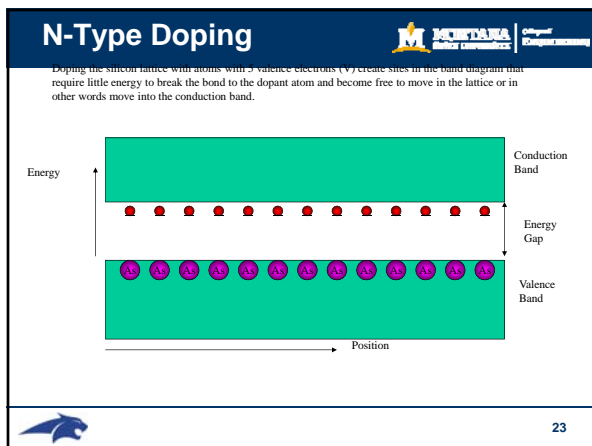
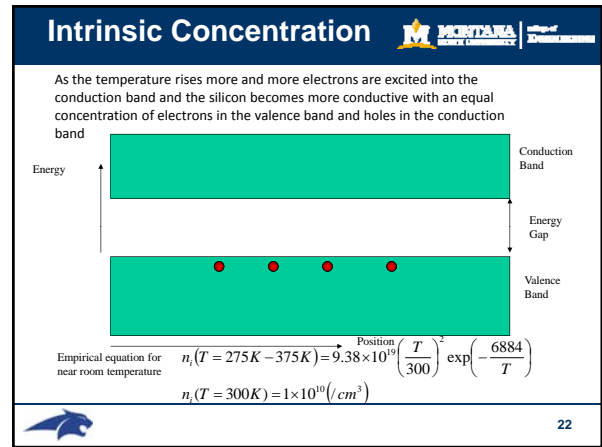
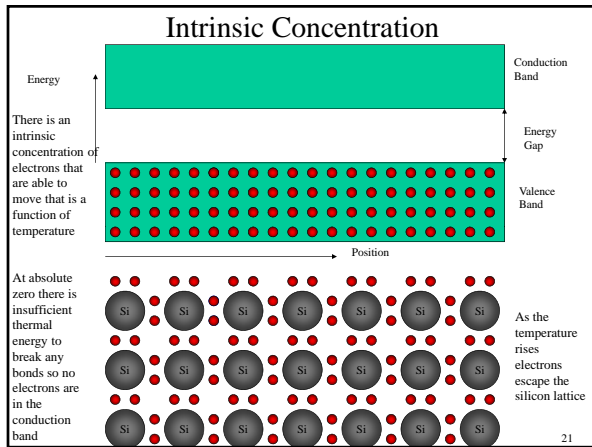
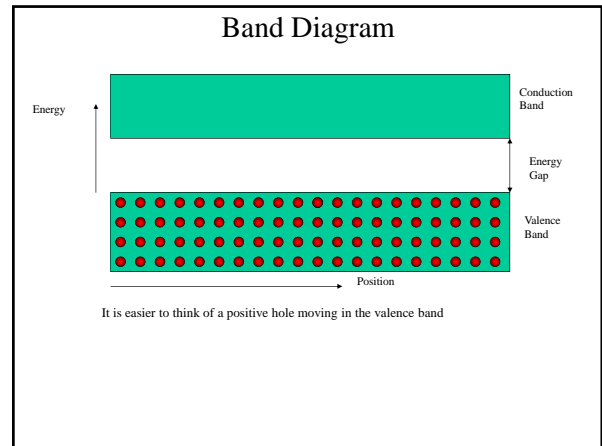
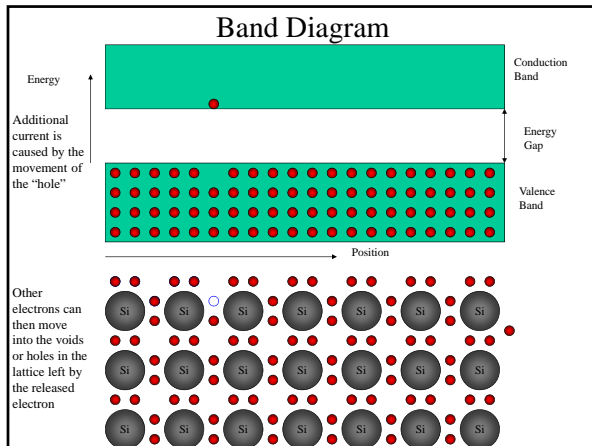
Valence Band

Position

This corresponds to an electron jumping from the valence band to the conduction band

Thermal energy causes an electron to break its bond to the silicon lattice

The electron is now free to move through the silicon lattice



P-Type Doping

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Doping the silicon lattice with atoms with 3 valence electrons (III) create sites in the band diagram that require little energy to trap an electron into the dopant atom. Holes are created in the valence band that are free to move.

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P type Material

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Majority Charge Carriers are Positive holes

Few electrons

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Equilibrium Carrier Concentration

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- Law of Mass Action
 - At equilibrium the product of the majority and minority carrier concentration is a constant
$$n_0 p_0 = n_i^2$$
 - n_0 is the equilibrium electron concentration
 - p_0 is the equilibrium hole concentration
 - n_i is the intrinsic concentration

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N-Type Carrier Concentration

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- At room temperature in an N-type material, the electron concentration is approximately the donor doping concentration

$$n_0 = N_D$$

$$p_0 = \frac{n_i^2}{N_D}$$

- N_D is the donor doping concentration

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P-Type Carrier Concentration

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- At room temperature in a P-type material, the hole concentration is approximately the acceptor doping concentration

$$p_0 = N_A$$

$$n_0 = \frac{n_i^2}{N_A}$$

- N_A is the acceptor doping concentration

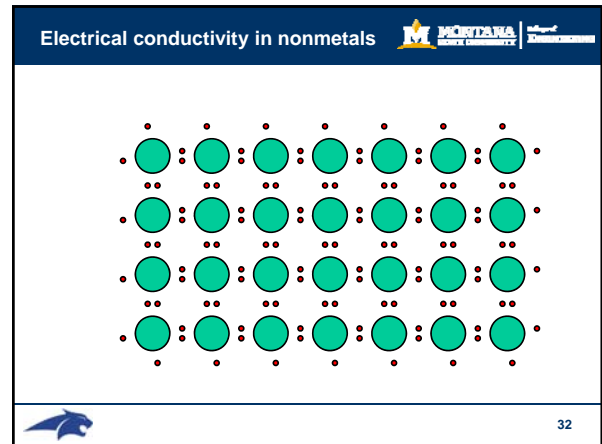
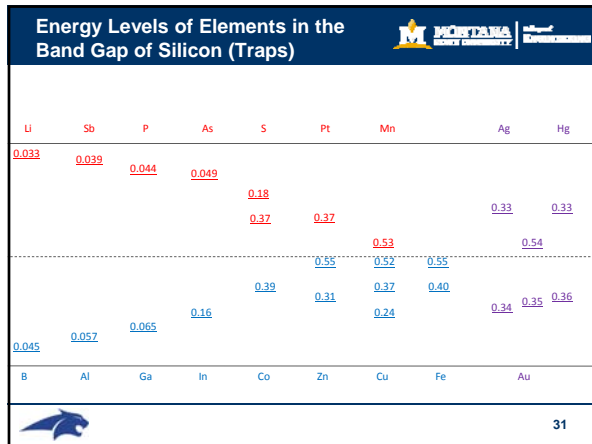
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Minority Concentration Dependence

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As the doping concentration increases the concentration of minority carriers decreases.

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Silicon

- Thermal energy breaks bond and creates a free electron and a hole
- The concentration of holes equals the concentration of free electrons
- An applied field causes electrons to contribute to electrical conduction
- Movement of electrons filling holes create net charge movement contributing to electrical conduction

Conductivity of Semiconductor

$$\sigma = ep\mu_h + en\mu_e \quad \mu_h < \mu_e$$

- p: concentration of holes
- n: concentration of electrons
- μ_h : hole mobility
- μ_e : electron mobility

$$\sigma = \frac{1}{\rho} \Rightarrow \rho = \frac{1}{\sigma}$$